**Nomenclature Lab**

**Directions for lab set-up**: Once your lab notebook is prepared, you will move around the lab in order to fill in the chart below. This is practice in writing and naming Ionic Compounds. Copying this information from other students does not provide you any practice for the upcoming test.

Purpose: Write the formula, name and description of Binary and Ternary Ionic Compounds.

Procedure: On each lab station there will be several vials. Move from station to station, recording the information needed to complete the data table below. Under the cation and anion you will be writing the symbol AND the charge. **You will need to crisscross down to write the formula**, or you will be writing the name.

Data: The Vials will be numbered **1A through 16A and 1B through 16B**.

Prepare your data table IN ADVANCE with 16 rows (1A to 16A) and the following columns:

 \*\*Also prepare data table #2 with 1B through 16B.

Be sure to show ALL CRISS CROSS WORK

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| Vial # | Description | Cation | Anion | Formula | Name |
| 1A |  |  |  |  |  |
| …. |  |  |  |  |  |
| 16A |  |  |  |  |  |

Conclusion Questions (only complete these once):

1. What is the difference between ionic compounds and covalent compounds?
2. How can you tell the difference between a multivalent cation and a binary ionic cation when looking at a formula (i.e. BeO vs NiCl2)?
3. Write down 2-3 properties of ionic compounds.
4. When looking at the vials, what trend do you see as to what is causing the compounds to have color other than white?